

# Previous Years Question Bank

## ONE MARK QUESTIONS

1. In electrolysis of water, why is the volume of gas collected over one electrode double that of gas collected over the other electrode.

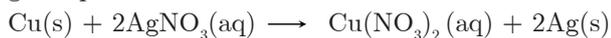
**Ans :** [CBSE S.R 2018, CBSE 2009, 2012, 2013]

Water contains hydrogen and oxygen in the ratio of 2 : 1, that is why volume of H<sub>2</sub> is double than that of oxygen.

2. What can be seen when a strip of copper metal is placed in a solution of silver nitrate?

**Ans :** [CBSE 2015]

The solution will become blue, shiny silver metal will get deposited.



3. State one industrial application of reduction process.

**Ans :** [CBSE 2015]

It is used in the extraction of metals e.g.,  
$$\text{ZnO(s)} + \text{C(s)} \xrightarrow{\text{Heat}} \text{Zn(s)} + \text{CO}_2(\text{g})$$

4. Which one of the following is a chemical change? Give reason also.

(a) Burning of wax (b) Melting of wax

**Ans :** [CBSE 2014]

Burning of wax is a chemical change because new products with new properties will be formed on burning.

5. Which one is a chemical change:  
Rusting of iron or melting of iron?

**Ans :** [CBSE 2014]

Rusting of iron is a chemical change.

6. State one basic difference between a physical change and a chemical change.

**Ans :** [CBSE 2014, 2011]

In a physical change, no new substance is formed. In a chemical change, new substance(s) with new properties is/are formed.

7. Name the oxidising and reducing agent in the following reaction:



**Ans :** [CBSE 2013]

CuO is oxidising agent H<sub>2</sub> is reducing agent.

8. Write a complete balanced chemical equation for the

following reaction:



**Ans :** [CBSE 2013]



9. Name and state the law which is kept in mind when we balance chemical equations.

**Ans :** [CBSE 2012]

Law of conservation of mass. It states **Matter can neither be created nor be destroyed**. The total mass of reactants must be equal to total mass of products.

10. What is meant by a chemical reaction?

**Ans :** [CBSE 2011]

Chemical reaction, a process in which one or more substances, the reactants, are converted to one or more different substances, the products. Substances are either chemical elements or compounds. A chemical reaction rearranges the constituent atoms of the reactants to create different substances as products

11. 
$$\text{AgNO}_3(\text{aq}) + \text{NaCl}(\text{aq}) \longrightarrow \text{AgCl} \downarrow + \text{NaNO}_3(\text{aq})$$
  
$$\text{FeS} + \text{H}_2\text{SO}_4 \longrightarrow \text{FeSO}_4 + \text{H}_2\text{S} \uparrow$$

Consider the above mentioned two chemical equations with two different kinds of arrows (↑ and ↓) along with the product. What do these two different arrows indicate?

**Ans :** [CBSE 2011]

↑ shows evolution of gas, ↓ represents a precipitate.

12. Hydrogen being a highly inflammable gas and oxygen being a supporter of combustion, yet water, a compound made up of hydrogen and oxygen is used to extinguish fire. Why?

**Ans :** [CBSE 2011]

H<sub>2</sub>O is a compound constituted of hydrogen of oxygen elements and being a compound it has different properties as compared to its constituting elements.

13. 
$$\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$$
, name the type of reaction.

**Ans :** [CBSE 2011]

Combination reaction.

14. If copper metal is heated over a flame it develops a coating. What is the colour and composition of coating?

**Ans :** [CBSE 2011]

Black coloured coating is formed. It is a due to formation of copper oxide.

15. Write a balanced chemical equation to represent the following reaction: carbon monoxide reacts with hydrogen gas at 340 atm to form methyl alcohol.  
**Ans :** [CBSE 2011]  

$$\text{CO(g)} + 2\text{H}_2\text{(g)} \xrightarrow{340 \text{ atm}} \text{CH}_3\text{OH(l)}$$
16. Which one is a chemical change: fermentation of fruit juice or diluting fruit juice?  
**Ans :** [CBSE 2011]  
 Fermentation of fruit juice is a chemical change.
17. Is burning of a candle wax a physical or a chemical change?  
**Ans :** [CBSE 2011]  
 Burning of a candle wax is a chemical change.
18. Write a balanced equation for the chemical reaction that can be characterised as precipitation reaction.  
**Ans :** [CBSE 2011]  

$$\text{AgNO}_3\text{(aq)} + \text{NaCl(aq)} \longrightarrow \text{AgCl(s)} + \text{NaNO}_3\text{(aq)}$$
 It is a precipitation reaction.
19. State the main difference between endothermic reaction and an exothermic reaction.  
**Ans :** [CBSE 2011]  
 In endothermic reaction, heat is absorbed. In exothermic reaction, heat is evolved.
20. What happens chemically when quick lime is added to water filled in a bucket?  
**Ans :** [CBSE 2010, 2008]  
 Calcium hydroxide (Slaked lime) is formed with evolution of heat and hissing sound.  

$$\text{CaO(s)} + \text{H}_2\text{O(l)} \longrightarrow \text{Ca(OH)}_2\text{(aq)}$$
21. Define oxidation and reduction.  
**Ans :** [CBSE 2011, 2010]  
 Oxidation is a process in which oxygen is added or loss of electrons take place. Reduction is a process in which hydrogen is added or gain of electrons takes place.
22. Give an example of double displacement reaction (only with complete balanced equation).  
**Ans :** [CBSE 2011, 2010]  

$$\text{BaCl}_2\text{(aq)} + \text{H}_2\text{SO}_4\text{(dil.)} \longrightarrow \text{BaSO}_4\text{(s)} + 2\text{HCl(aq)}$$
 It is a double displacement reaction.
23. On what basis is a chemical reaction balanced?  
**Ans :** [CBSE 2010, 2008]  
 Chemical equation is balanced on the basis of law of conservation of mass.
24. What change in colour is observed when white silver chloride is left exposed to sun-light? State the type of chemical reaction in this change.  
**Ans :** [CBSE 2010, 2009]  
 Grey coloured silver metal is formed and pungent smelling chlorine gas is evolved.  

$$2\text{AgCl(s)} \xrightarrow{\text{sunlight}} 2\text{Ag(s)} + \text{Cl}_2\text{(g)}$$
- It is photochemical decomposition reaction.
25. Write a balanced chemical equation for the reaction between sodium chloride and silver nitrate indicating the physical state of the reactants and the products.  
**Ans :** [CBSE 2010]  

$$\underset{\text{(Silver Nitrate)}}{\text{AgNO}_3\text{(aq)}} + \underset{\text{(Sodium Chloride)}}{\text{NaCl(aq)}} \xrightarrow{\text{Heat}} \underset{\text{(Silver Chloride)}}{\text{AgCl(s)}} + \underset{\text{(Sodium Nitrate)}}{\text{NaNO}_3\text{(aq)}}$$
26. Complete and balance the following equation:  

$$\text{Fe}_2\text{O}_3 + \text{Al} \longrightarrow$$
  
**Ans :** [CBSE 2010]  

$$2\text{Fe}_2\text{O}_3 + 2\text{Al} \longrightarrow 2\text{Fe} + \text{Al}_2\text{O}_3$$
27. Balance the following chemical equation:  

$$\text{Pb(NO}_3)_2 \longrightarrow \text{PbO} + \text{NO}_2 + \text{O}_2$$
  
**Ans :** [CBSE 2009]  

$$2\text{Pb(NO}_3)_2 \xrightarrow{\text{Heat}} 2\text{PbO} + 4\text{NO}_2 + \text{O}_2$$
28. Identify the type of reaction in the following example:  

$$\text{Na}_2\text{SO}_4\text{(aq)} + \text{BaCl}_2\text{(aq)} \longrightarrow \text{BaSO}_4\text{(s)} + 2\text{NaCl(aq)}$$
  
**Ans :** [CBSE 2008(C)]  
 Double displacement reaction.
29. Identify the type of reaction in the following example:  

$$\text{Fe(s)} + \text{CuSO}_4\text{(aq)} \longrightarrow \text{FeSO}_4\text{(aq)} + \text{Cu(s)}$$
  
**Ans :** [CBSE 2008(C)]  
 Displacement reaction.
30. Identify the type of reaction in the following example:  

$$2\text{H}_2\text{(g)} + \text{O}_2\text{(g)} \longrightarrow 2\text{H}_2\text{O(l)}$$
  
**Ans :** [CBSE 2008(C)]  
 Combination reaction.
31. Balance the given chemical equation:  

$$\text{Al(s)} + \text{CuCl}_2\text{(aq)} \longrightarrow \text{AlCl}_3\text{(aq)} + \text{Cu(s)}$$
  
**Ans :** [CBSE 2008(C)]  

$$2\text{Al(s)} + 3\text{CuCl}_2\text{(aq)} \longrightarrow 2\text{AlCl}_3\text{(aq)} + 3\text{Cu(s)}$$
32. Balance the given chemical equation:  

$$\text{FeSO}_4\text{(s)} \xrightarrow{\text{Heat}} \text{Fe}_2\text{O}_3\text{(s)} + \text{SO}_2\text{(g)} + \text{SO}_3\text{(g)}$$
  
**Ans :** [CBSE 2008]  

$$2\text{FeSO}_4\text{(s)} \xrightarrow{\text{Heat}} \text{Fe}_2\text{O}_3\text{(s)} + \text{SO}_2\text{(g)} + \text{SO}_3\text{(g)}$$
33. Balance the following chemical equation:  

$$\text{Fe(s)} + \text{H}_2\text{O(g)} \longrightarrow \text{Fe}_3\text{O}_4\text{(s)} + \text{H}_2\text{(g)}$$
  
**Ans :**  

$$3\text{Fe(s)} + 4\text{H}_2\text{O(g)} \longrightarrow \text{Fe}_3\text{O}_4\text{(s)} + 4\text{H}_2\text{(g)}$$
34. On adding dilute hydrochloric acid to copper oxide powder, the solution formed is blue green. Predict the new compound formed which imparts a blue green colour to the solution.  
**Ans :** [CBSE 2008]  

$$\text{CuO} + 2\text{HCl} \longrightarrow \text{CuCl}_2 + 2\text{H}_2\text{O}$$
 Copper chloride solution imparts blue green colour to the solution.
35. Why is respiration considered as exothermic process?  
**Ans :** [CBSE 2008]





They become dirty white on heating.

- b. Ferric oxide, sulphur dioxide and sulphur trioxide are formed:



It is a decomposition reaction.

59. Reaction of compound X with aluminium is used to join railway tracks or cracked machine parts.
- Identify the compound.
  - Name the reaction.
  - Write a balanced chemical equation for the reaction.

**Ans :** [CBSE 2012]

- The compound X is  $\text{Fe}_2\text{O}_3$  (Ferric oxide) or Iron (III) oxide.
- It is called Thermite Reaction.
- $2\text{Al}(\text{s}) + \text{Fe}_2\text{O}_3(\text{s}) \xrightarrow{\text{Heat}} \text{Al}_2\text{O}_3(\text{s}) + 2\text{Fe}(\text{l})$

60. Using balanced chemical equation explain the difference between a displacement reaction and a double displacement reaction.

**Ans :** [CBSE 2012, 2011]

**Displacement reaction:** A reaction in which a more reactive element displaces a less reactive element from its salt solution e.g.,



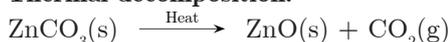
**Double displacement reaction:** A reaction in which two compounds exchange their ions to form two new compounds e.g.,



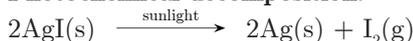
61. Give an example each for thermal decomposition and photochemical decomposition reactions. Write balanced chemical equation also.

**Ans :** [CBSE 2012]

**Thermal decomposition:**



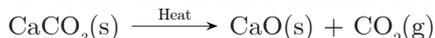
**Photochemical decomposition:**



62. Why are decomposition reactions called the opposite of combination reactions? Give chemical equations for these reactions.

**Ans :** [CBSE 2012]

In decomposition reactions, a compound is broken down into one or more elements or compounds e.g.,



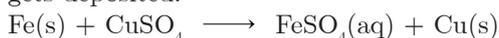
In combination reactions, two or more elements or compounds combine to form a new compound. Therefore, decomposition reactions are opposite to combination reactions e.g.,



63. Why does the colour of copper sulphate solution changes when an iron pin is dipped in it?

**Ans :** [CBSE 2012]

Iron being more reactive displaces copper from copper sulphate (Blue) solution to form iron (II) sulphate (Pale green) solution and reddish brown copper metal gets deposited.



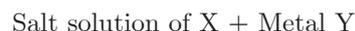
64. Translate the following statements into chemical equations and then balance them:

- Hydrogen gas combines with nitrogen to form ammonia.
- Hydrogen sulphide gas burns in air to give water and sulphur dioxide.
- Barium chloride reacts with aluminium sulphate to give aluminium chloride and a precipitate of barium sulphate. State the two types in which this reaction can be classified.
- Potassium reacts with water to give potassium hydroxide and hydrogen gas.

**Ans :** [CBSE 2012]

- $3\text{H}_2(\text{g}) + \text{N}_2(\text{g}) \longrightarrow 2\text{NH}_3(\text{g})$
- $2\text{H}_2\text{S}(\text{g}) + 3\text{O}_2(\text{g}) \longrightarrow 2\text{H}_2\text{O}(\text{l}) + 2\text{SO}_2(\text{g})$
- $3\text{BaCl}_2(\text{aq}) + \text{Al}_2(\text{SO}_4)_3(\text{aq}) \longrightarrow 3\text{BaSO}_4(\text{s}) + 2\text{AlCl}_3(\text{aq})$   
This reaction can be classified as a double displacement reaction as the two reacting species ( $\text{BaSO}_4$ ,  $\text{Al}_2\text{SO}_4$ )<sub>3</sub> undergoes mutual exchange of ions. It can also be classified as precipitation reaction, since a white ppt. of  $\text{BaSO}_4$  is obtained.
- $2\text{K}(\text{s}) + 2\text{H}_2\text{O}(\text{l}) \longrightarrow 2\text{KOH}(\text{aq}) + \text{H}_2(\text{g})$

65. When a metal X is added to salt solution of metal Y, the following chemical reaction takes place:



Mention the inference you draw regarding the reactivity of metal X and Y and also the type of reaction. State the reason of your conclusions.

**Ans :** [CBSE 2012]

X is more reactive than Y because, X displaces Y from its salt solution. Hence, this reaction is an example of displacement reaction.

66. Identify the type of each of the following reactions:

- A reaction in which a single product is formed from two or more reactants.
- The reaction mixture becomes warm.
- An insoluble substance is formed.
- External surface of the container in which reaction takes place becomes cold.

**Ans :** [CBSE 2012]

- Combination reaction
- Exothermic reaction
- Precipitation reaction (Double displacement reaction)
- Endothermic reaction

67. A solution of potassium chloride when mixed with silver nitrate solution, an insoluble white substance is formed. Write the chemical reaction involved and also mention the type of reaction.

**Ans :** [CBSE 2012, 2011, 2010]



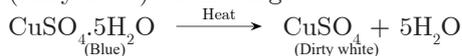
It is a double displacement reaction.

68. Using a suitable chemical equation justify that some chemical reactions are determined by

- Change in colour.
- Change in temperature.

**Ans :** [CBSE 2011]

- a. **Change in colour:** Blue coloured crystalline copper sulphate crystals changes to white  $\text{CuSO}_4$  (anhydrous) on heating.



- b. **Change in temperature:**  
 $\text{CaO(s)} + \text{H}_2\text{O(l)} \longrightarrow \text{Ca(OH)}_2\text{(aq)} + \text{Heat}$

69. (a) A solution of substance X is used for white washing. What is substance X? Write the chemical reaction of X with water.  
(b) Why does the colour of copper sulphate solution changes when iron nail is dipped in it?

**Ans :** [CBSE 2011]

- a. X is CaO, calcium oxide  
 $\text{CaO(s)} + \text{H}_2\text{O(l)} \longrightarrow \text{Ca(OH)}_2\text{(aq)}$   
b. It is because Fe displaces Cu from  $\text{CuSO}_4$  (blue) solution to form  $\text{FeSO}_4$  (pale green) and reddish brown Cu metal gets deposited.

70. Balance the following reactions:

- a.  $\text{BaCl}_2 + \text{H}_2\text{SO}_4 \longrightarrow \text{BaSO}_4 + \text{HCl}$   
b.  $\text{Ca(OH)}_2 + \text{HNO}_3 \longrightarrow \text{Ca(NO}_3)_2 + \text{H}_2\text{O}$   
c.  $\text{Pb(NO}_3)_2 \longrightarrow \text{PbO} + \text{NO}_2 + \text{O}_2$   
d.  $\text{MnO}_2 + \text{HCl} \longrightarrow \text{MnCl}_2 + \text{Cl}_2 + \text{H}_2\text{O}$

**Ans :** [CBSE 2011, 2009, 2008, 2008C]

- a.  $\text{BaCl}_2\text{(aq)} + \text{H}_2\text{SO}_4\text{(dil)} \longrightarrow \text{BaSO}_4\text{(s)} + 2\text{HCl(aq)}$   
b.  $\text{Ca(OH)}_2\text{(aq)} + 2\text{HNO}_3 \longrightarrow \text{Ca(NO}_3)_2\text{(aq)} + 2\text{H}_2\text{O(l)}$   
c.  $2\text{Pb(NO}_3)_2\text{(s)} \xrightarrow{\text{Heat}} 2\text{PbO(s)} + 4\text{NO}_2\text{(g)} + \text{O}_2\text{(g)}$   
d.  $\text{MnO}_2 + 4\text{HCl} \longrightarrow \text{MnCl}_2 + \text{Cl}_2 + 2\text{H}_2\text{O}$

71. Write the balanced equation for the following reactions and identify the type of reaction in each case:

- a. Potassium bromide + Barium iodide  $\longrightarrow$   
Barium bromide + Potassium Iodide  
b. Hydrogen(g) + Chlorine(g)  $\longrightarrow$  Hydrogen chloride(g)

**Ans :** [CBSE 2011]

- a.  $2\text{KBr(aq)} + \text{BaI}_2 \longrightarrow \text{BaBr}_2\text{(aq)} + 2\text{KI(aq)}$   
It is double displacement reaction.  
b.  $\text{H}_2\text{(g)} + \text{Cl}_2\text{(g)} \xrightarrow{\text{sunlight}} 2\text{HCl(g)}$   
It is combination reaction.

72. A zinc plate was put into solution of copper sulphate kept in a glass container. It was found that blue colour of the solution gets fader and fader with passage of time. After few days when zinc plate was taken out of the solution, a number of holes were observed on it.

- a. State the reason for the changes observed on zinc plate.  
b. Write the chemical equation for the reaction involved.

**Ans :** [CBSE 2011]

- a. Zinc displaces copper from copper sulphate solution to form colourless  $\text{ZnSO}_4$  and copper metal is deposited. Zinc gets consumed due to which holes are formed.  
b.  $\text{Zn(s)} + \text{CuSO}_4\text{(aq)} \longrightarrow \text{ZnSO}_4\text{(aq)} + \text{Cu(s)}$

73. A white salt on heating decomposes to give brown

fumes and a white residue is left behind.

- a. Name the salt.  
b. Write the equation for the decomposition reaction.

**Ans :** [CBSE 2011]

- a. Lead nitrate  
b.  $2\text{Pb(NO}_3)_3\text{(s)} \xrightarrow{\text{Heat}} 2\text{PbO(s)} + 4\text{NO}_2\text{(g)} + \text{O}_2\text{(g)}$   
(White) (Residue) (Brown)

74. When a solution of potassium iodide is added to a solution of lead nitrate in a test tube, a reaction takes place.

- a. What type of reaction is this?  
b. Write the balanced chemical equation to represent the above reaction.

**Ans :** [CBSE 2011, 2010, 2008]

- a. It is precipitation as well as double displacement reaction.  
b.  $\text{Pb(NO}_3)_2\text{(aq)} + 2\text{KI(aq)} \longrightarrow \text{PbI}_2\text{(s)} + 2\text{KNO}_3\text{(aq)}$

75. Define combination reaction. Give one example of a combination reaction which is also exothermic.

**Ans :** [CBSE 2011]

**Combination reaction:** The reaction in which two or more elements or compounds combine to form compound(s) e.g.,

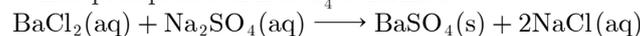


It is an exothermic, combination reaction.

76. What happens when an aqueous solution of sodium sulphate reacts with an aqueous solution of barium chloride? State the physical conditions of reactants in which the reaction between them will not take place. Write the balanced chemical equation for the reaction and name the type of reaction.

**Ans :** [CBSE 2016, 2010]

White precipitate of  $\text{BaSO}_4$  is formed.



It is a double displacement reaction. If reactants are taken in solid state, products will not be formed.

77. What is redox reaction? When a magnesium ribbon burns in air with a dazzling flame and forms a white ash, is magnesium oxidised or reduced. Why?

**Ans :** [CBSE 2010, 2009]

Redox reaction is a reaction in which oxidation and reduction takes place simultaneously.

Mg is getting oxidised because it is gaining oxygen to form magnesium oxide.

78. Write any two observations in an activity which may suggest that a chemical reaction has taken place. Give an example in support of your answer.

**Ans :** [CBSE 2010]

- a. **Change in colour:**  
 $\text{AgNO}_3\text{(aq)} + \text{KI(aq)} \longrightarrow \text{AgI(s)} + \text{KNO}_3\text{(aq)}$   
(Yellow ppt)

- b. **Evolution of gas:**  
 $\text{Zn(s)} + \text{H}_2\text{SO}_4\text{(aq)} \longrightarrow \text{ZnSO}_4\text{(aq)} + \text{H}_2\text{(g)}$

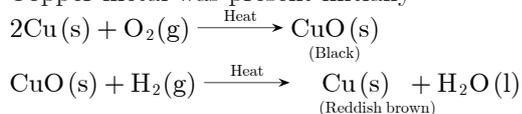
79. When the powder of common metal is heated in open china dish its colour turns black. However, when hydrogen is passed over the hot black substance so formed, it regains its original colour. Based on the

above information answer the following questions:

- What type of chemical reaction takes place in each of the two given steps?
- Name the metal initially taken in powdered form. Write balanced equations for both the reactions.

**Ans :** [CBSE 2010]

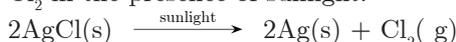
- Oxidation reaction, Redox reaction
- Copper metal was present initially



- 80.** Why do we store silver chloride in dark coloured bottle? Explain in brief.

**Ans :** [CBSE 2010]

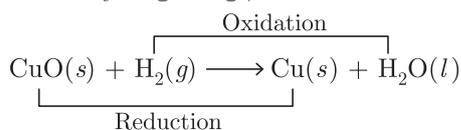
It is done so as to cut off the exposure to sunlight. AgCl is photosensitive, it will decompose to Ag and Cl<sub>2</sub> in the presence of sunlight.



- 81.** "Oxidation and reduction processes occur simultaneously". Justify this statement with the help of example.

**Ans :** [CBSE 2010]

Oxidation involves loss of electrons or addition of oxygen. Whereas reduction involves gain of electrons or addition of hydrogen e.g.,



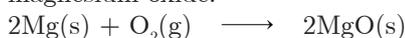
H<sub>2</sub> is getting oxidised to H<sub>2</sub>O, CuO is getting reduced to Cu.

It shows oxidation and reduction occur simultaneously.

- 82.** When magnesium ribbon burns in air or oxygen, a product is formed. State the type of chemical reaction and name the product formed in the reaction. Write balanced chemical equation for the reaction.

**Ans :** [CBSE 2009(C)]

It is a combination reaction. The product formed is magnesium oxide:



- 83.** Distinguish between a displacement reaction and a double displacement reaction. Identify the displacement and the double displacement reaction from the following reactions:

- $\text{HCl(aq)} + \text{NaOH(aq)} \longrightarrow \text{NaCl(aq)} + \text{H}_2\text{O(l)}$
- $\text{Fe(s)} + \text{CuSO}_4(\text{aq}) \longrightarrow \text{FeSO}_4(\text{aq}) + \text{Cu(s)}$

**Ans :** [CBSE 2009]

**Displacement Reaction:** When a more reactive metal displaces a less reactive metal from its salt solution.

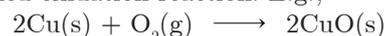
**Double displacement reaction:** When two compounds exchange their ions to form two new compounds.

- Double displacement reaction.
- Displacement reaction.

- 84.** What is an oxidation reaction? Give an example of oxidation reaction. Is oxidation an exothermic or an endothermic reaction?

**Ans :** [CBSE 2009]

**Oxidation reaction:** The reaction in which O<sub>2</sub> is added or H<sub>2</sub> is removed or loss of electrons takes place is called oxidation reaction. E.g.,



It is an exothermic reaction.

## THREE MARKS QUESTIONS

- 85.** Decomposition reactions require energy either in the form of heat, light or electricity for breaking down the reactants. Write an equation each for decomposition reactions where energy is supplied in the form of heat, light and electricity.

**Ans :** [CBSE 2018, 2014]

- $\text{CaCO}_3(\text{s}) \xrightarrow{\text{Heat}} \text{CaO(s)} + \text{CO}_2(\text{g})$
- $2\text{AgBr(s)} \xrightarrow{\text{sunlight}} 2\text{Ag(s)} + \text{Br}_2(\text{g})$
- $2\text{H}_2\text{O(l)} \xrightarrow{\text{Electricity}} 2\text{H}_2(\text{g}) + \text{O}_2(\text{g})$

- 86.** In the electrolysis of water:

- Name the gas collected at the cathode and anode respectively.
- Why is volume of gas collected at one electrode double than that at the other? Name this gas.
- How will you test this gas?

**Ans :** [CBSE 2012, CBSE Sample Paper 2018]

- Hydrogen is collected at the cathode, oxygen is collected at the anode.
- It is because H<sub>2</sub>O contains hydrogen and oxygen in the ratio 2 : 1.
- Bring a burning matchstick near the gas, if the gas burns with 'pop' sound, the gas is H<sub>2</sub>.

- 87.** Define the term decomposition reaction. Give one example each of thermal decomposition and electrolytic decomposition reactions.

**Ans :** [CBSE 2016]

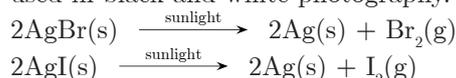
**Decomposition reaction:** The reaction in which a compound is broken down into simpler elements or compounds:

- $\text{CaCO}_3(\text{s}) \xrightarrow{\text{Heat}} \text{CaO(s)} + \text{CO}_2(\text{g})$
- $2\text{H}_2\text{O(l)} \xrightarrow{\text{Electricity}} 2\text{H}_2(\text{g}) + \text{O}_2(\text{g})$

- 88.** Name two salts that are used in black and white photography. Give equations for the reactions when these are exposed to sunlight.

**Ans :** [CBSE 2016]

AgBr (Silver bromide) and AgI (Silver iodide) are used in black and white photography.



- 89.** State one example each characterised by following along with suitable chemical equation.

- Change in state,
- Evolution of gas,
- Change in temperature.

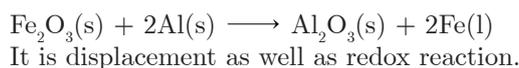
**Ans :** [CBSE 2016]

- Change in state:**  
 $\text{AgNO}_3(\text{aq}) + \text{HCl(aq)} \longrightarrow \text{AgCl(s)} + \text{HNO}_3(\text{aq})$
- Evolution of gas:**



reaction and identify the type of reaction.  
Thermite reaction, iron (III) oxide reacts with aluminium and give molten iron and aluminium oxide.

**Ans :** [CBSE 2012]



99. An aqueous solution of metal nitrate 'P' reacts with sodium bromide solution to form yellow precipitate 'Q' which is used in photography. 'Q' on exposure to sunlight undergoes decomposition to form metal present along with a reddish brown gas. Identify 'P' and 'Q' write the balanced chemical equation for the chemical reaction. List the two categories in which reaction can be placed.

**Ans :** [CBSE 2012]

'P' is silver nitrate ( $\text{AgNO}_3$ )  

$$\text{AgNO}_3(\text{aq}) + \text{NaBr}(\text{aq}) \longrightarrow \underset{\text{(P)}}{\text{AgBr}(\text{s})} + \underset{\text{(Yellow ppt)}}{\text{NaNO}_3(\text{aq})}$$
 This reaction is categorised as double displacement as well as precipitation reaction.  

$$2\text{AgBr}(\text{s}) \xrightarrow{\text{sunlight}} 2\text{Ag}(\text{s}) + \text{Br}_2(\text{g})$$
 The above reaction is photochemical decomposition reaction.

100. A green coloured hydrated metallic salt on heating loses water of crystallisation molecules and gives a gas with suffocating smell. Identify the salt and write the chemical equation for the reaction.

**Ans :** [CBSE 2011]

The salt is  $\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$  (Hydrated ferrous sulphate)  

$$\text{FeSO}_4 \cdot 7\text{H}_2\text{O} \xrightarrow{\text{Heat}} \text{FeSO}_4(\text{s}) + 7\text{H}_2\text{O}(\text{l})$$

$$2\text{FeSO}_4(\text{s}) \xrightarrow{\text{Heat}} \text{Fe}_2\text{O}_3(\text{s}) + \text{SO}_2(\text{g}) + \text{SO}_3(\text{g})$$

101. a. Can combination reaction be an oxidation reaction?  
 b. How will you test whether the gas evolved in a reaction is hydrogen?  
 c. Why does copper not evolve hydrogen on reacting with dilute sulphuric acid?

**Ans :** [CBSE 2011]

- a. Yes, combination reaction can be called a oxidation reaction.  
 b. Bring a burning splinter near the gas, if it burns with pop sound, it is hydrogen gas.  
 c. It is because copper is less reactive than hydrogen.

102. Write balanced equations for the following reactions:  
 a. Aluminium + Bromine  $\longrightarrow$  Aluminium bromide  
 b. Calcium carbonate  $\xrightarrow{\text{Heat}}$

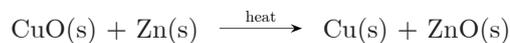
Calcium oxide + Carbon dioxide

- c. Silver chloride  $\xrightarrow{\text{sunlight}}$  Silver + Chlorine

**Ans :** [CBSE 2011]

- a.  $2\text{Al}(\text{s}) + 3\text{Br}_2(\text{g}) \longrightarrow 2\text{AlBr}_3(\text{s})$   
 b.  $\text{CaCO}_3(\text{s}) \xrightarrow{\text{Heat}} \text{CaO}(\text{s}) + \text{CO}_2(\text{g})$   
 c.  $2\text{AgCl}(\text{s}) \xrightarrow{\text{sunlight}} 2\text{Ag}(\text{s}) + \text{Cl}_2(\text{g})$

103. a. Why is respiration considered an exothermic reaction?  
 b. Define the terms oxidation and reduction,  
 c. Identify the substance that is oxidised and reduced in the following reaction:



**Ans :** [CBSE 2011]

- a. In respiration, energy is released, therefore, it is considered an exothermic reaction.  
 b. Oxidation involves addition of  $\text{O}_2$  or loss of electrons. Reduction involves addition of  $\text{H}_2$  or gain of electrons.  
 c. Zn is getting oxidised, CuO is getting reduced.

104. Write balanced chemical equations for the following reactions:

- a. Hydrogen sulphide gas burns in air to give water and sulphur dioxide.  
 b. Barium chloride reacts with zinc sulphate to give zinc chloride and barium sulphate.  
 c. Natural gas burns in air to form carbon dioxide and water.

**Ans :** [CBSE 2011]

- a.  $2\text{H}_2\text{S}(\text{g}) + 3\text{O}_2(\text{g}) \longrightarrow 2\text{H}_2\text{O}(\text{l}) + 2\text{SO}_2(\text{g})$   
 b.  $\text{BaCl}_2(\text{aq}) + \text{ZnSO}_4(\text{aq}) \longrightarrow \text{BaSO}_4(\text{s}) + \text{ZnCl}_2(\text{aq})$   
 c.  $\text{CH}_4(\text{g}) + 2\text{O}_2 \longrightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{l})$

## FIVE MARKS QUESTIONS

105. Write balanced chemical equation for the following statements:

- a. NaOH solution is heated with zinc granules.  
 b. Excess of carbon dioxide is passed through lime water.  
 c. Dilute sulphuric acid is added to sodium carbonate.  
 d. Egg shell is dropped in hydrochloric acid,  
 e. Copper (II) oxide reacts with dilute hydrochloric acid.

**Ans :** [CBSE 2016]

- a.  $\text{Zn}(\text{s}) + 2\text{NaOH} \longrightarrow \text{Na}_2\text{ZnO}_2 + \text{H}_2$   
 b.  $\text{Ca}(\text{OH})_2 + 2\text{CO}_2 \longrightarrow \text{Ca}(\text{HCO}_3)_2$   
 c.  $\text{Na}_2\text{CO}_3 + \text{H}_2\text{SO}_4 \longrightarrow \text{Na}_2\text{SO}_4 + \text{H}_2\text{O} + \text{CO}_2$   
 d.  $\text{CaCO}_3(\text{s}) + 2\text{HCl}(\text{dil}) \longrightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$   
 e.  $\text{CuO}(\text{s}) + 2\text{HCl} \longrightarrow \text{CuCl}_2 + \text{H}_2\text{O}$

106. Identify the type of chemical reaction in the following statement and define each of them:

- a. Digestion of food in our body.  
 b. Rusting of iron.  
 c. Heating of manganese dioxide with aluminium powder.  
 d. Blue colour of copper sulphate solution disappears when iron filings are added to it.  
 e. Dilute hydrochloric acid is added to sodium hydroxide solution to form sodium chloride and water.

**Ans :** [CBSE 2016]

- a. **Decomposition reaction:** It is a process in which a compound is broken down into simple substances.  
 b. **Oxidation:** The process in which oxygen is added or electrons are lost.  
 c. **Displacement reaction:** The reaction in which a more reactive element can displace a less reactive element from its salt solution. Oxidation and Reduction are taking place simultaneously in rusting of iron.

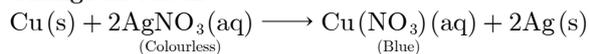
- d. **Displacement reaction:** The reaction in which a more reactive element can displace a less reactive element.
- e. **Neutralisation reaction:** The reaction in which acid reacts with base to form salt and water.

**107.** Define chemical reaction. State four observations which helps to determine whether a chemical reaction has taken place or not. Write one example of each observation with a balanced chemical equation.

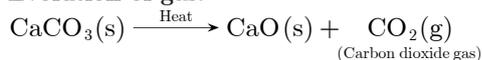
**Ans :** [CBSE 2015]

Chemical reaction is a reaction which represents a chemical change.

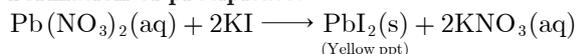
a. **Change in colour:**



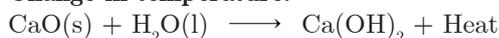
b. **Evolution of gas:**



c. **Formation of precipitate:**



d. **Change in temperature:**



- 108.** a. Define a balanced chemical equation. Why should an equation be balanced?
- b. Write a balanced chemical equation for the following reactions:
- (i) Phosphorus burns in the presence of chlorine to form phosphorus pentachloride.
- (ii) Burning of natural gas.
- (iii) The process of respiration.

**Ans :** [CBSE 2015]

a. Balanced chemical equation is a equation in which number of atoms of various elements are equal on both sides of the equation.

The equation should be balanced due to law of conservation of mass.

- b. (i)  $2\text{P(s)} + 5\text{Cl}_2(\text{g}) \longrightarrow 2\text{PCl}_5(\text{s})$
- (ii)  $\text{CH}_4(\text{g}) + \text{O}_2(\text{g}) \longrightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O(l)}$
- (iii)  $\text{C}_6\text{H}_{12}\text{O}_6(\text{s}) + 6\text{O}_2(\text{g}) \longrightarrow 6\text{CO}_2(\text{g}) + 6\text{H}_2\text{O(l)}$

- 109.** (a) Write one example for each of decomposition reaction carried out with the help of (i) Electricity, (ii) Heat, (iii) Light.
- (b) Which of the following statement is correct and why?

Copper can displace silver from silver nitrate solution and silver can displace copper from copper sulphate solution.

**Ans :** [CBSE 2014]

- a. (i)  $\text{CaCO}_3(\text{s}) \xrightarrow{\text{Heat}} \text{CaO(s)} + \text{CO}_2(\text{g})$
- (ii)  $2\text{AgBr(s)} \xrightarrow{\text{sunlight}} 2\text{Ag(s)} + \text{Br}_2(\text{g})$
- (iii)  $2\text{H}_2\text{O(l)} \xrightarrow{\text{Electricity}} 2\text{H}_2(\text{g}) + \text{O}_2(\text{g})$
- b. Copper can displace Ag from  $\text{AgNO}_3(\text{aq})$  solution because Cu is more reactive than Ag.
- $$\text{Cu(s)} + 2\text{AgNO}_3(\text{aq}) \longrightarrow \text{Cu(NO}_3)_2(\text{aq}) + 2\text{Ag(s)}$$

- 110.** Identify the type of reactions taking place in each of the following:
- a. Barium chloride solution is mixed with copper sulphate solution and white precipitate is formed.
- b. On heating copper powder in china dish, the

surface of copper powder turns black.

- c. On heating green coloured ferrous sulphate crystals, reddish brown solid is left and smell of a gas having odour of burning sulphur is experienced.
- d. Iron nails when left dipped in blue copper sulphate solution become reddish brown in colour and the blue colour of copper sulphate fades away.
- e. Quick lime reacts vigorously with water releasing a large amount of heat.

**Ans :** [CBSE Sample Paper 2009]

- a. Double displacement reaction,
- b. Oxidation,
- c. Decomposition reaction,
- d. Displacement reaction,
- e. Combination reaction.